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Unexpected Trends in the Stability and Dissociation Kinetics of Lanthanide(III) Complexes with Cyclen-Based Ligands Across the Lanthanide Series

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ABSTRACT: We report a detailed study of the thermodynamic stability and dissociation kinetics of lanthanide complexes with two ligands containing a cyclen unit, a methyl group, a picolinate arm and two acetate pendant arms linked to two nitrogen atoms of the macrocycle either in *cis* (1,4-H₃DO2APA) or *trans* (1,7-H₃DO2APA) positions. The stability constants of the Gd³⁺ complexes with these two ligands are very similar, with logK_{GdL} values of 16.98 and 16.33 for the complexes of 1,4-H₃DO2APA and 1,7-H₃DO2APA, respectively. The stability constants of the complexes with 1,4-H₃DO2APA follow the usual trend, increasing from logK_{LaL} = 15.96 to logK_{LuL} = 19.21. However, the stability of [Ln(1,7-DO2APA)] complexes decreases from logK = 16.33 for Gd³⁺ to 14.24 for Lu³⁺. The acid-catalyzed dissociation rates of the Gd³⁺ complexes differ by a factor of ~15, with rate constants (*k*₁), of 1.42 and 23.5 M⁻¹s⁻¹ for [Gd(1,4-DO2APA)] and [Gd(1,7-DO2APA)]. This difference is magnified across the lanthanide series to reach a five orders of magnitude higher *k*₁ for [Yb(1,7-DO2APA)] (1475 M⁻¹s⁻¹) than for [Yb(1,4-DO2APA)] (5.79 × 10⁻³ M⁻¹s⁻¹). The acid-catalysed mechanism involves the protonation of a carboxylate group, followed by a cascade of proton-transfer events that result in the protonation of a nitrogen atom of the cyclen unit. DFT calculations suggest a correlation between the strength of the Ln-O_{carboxylate} bonds and the kinetic inertness of the complex, with stronger bonds providing more inert complexes. The ¹H NMR resonance of the coordinated water molecule in the [Yb(1,7-DO2APA)] complex at 176 ppm provides a sizeable chemical exchange saturation transfer (CEST) effect thanks to a slow water exchange rate (15.9 ± 1.6) × 10³ s⁻¹.

INTRODUCTION

Lanthanide complexes based on the macrocyclic 1,4,7,10-tetraazacyclododecane (cyclen) containing different types and numbers of pendant arms were intensively investigated in the last decades due to the important medical and bioanalytical applications of their metal complexes.^{1,2} For example, gadolinium complexes of cyclen derivatives are currently used as contrast agents in magnetic resonance imaging (MRI),³ while analogous complexes with other lanthanide ions are currently being investigated as paramagnetic ¹H chemical exchange saturation transfer (paraCEST),⁴⁻⁶ ¹⁹F⁷⁻⁹ and paraSHIFT^{10,11} MRI contrast agents. Furthermore, many lanthanide cyclen-

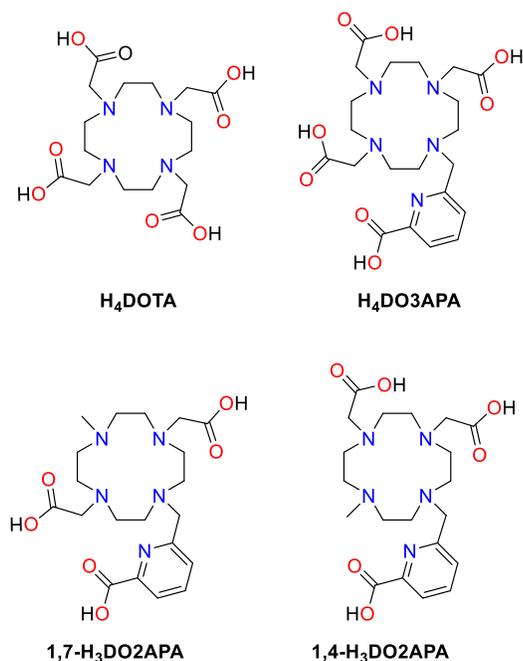
based luminescent complexes have been developed for application as optical probes in optical imaging or in cellulose studies.¹²⁻¹⁵

Whatever the targeted medical or biomedical application of a lanthanide complex, the chelate must present a high stability to avoid the toxic effects associated to the dissociation of the complex *in vivo*.^{16,17} Nowadays it is widely recognised that a high kinetic inertness with respect to their dissociation is more important than a high thermodynamic stability of the complex.^{18,19} For instance, it has been shown that lanthanide complexes with cyclen derivatives containing four acetamide pendant arms are extremely inert,²⁰ which allowed their safe *in vivo* application,²¹⁻²³ in spite of their low thermodynamic

stabilities compared with those of the carboxylate analogue DOTA.²⁰

We recently reported a series of ligands based on the cyclen platform functionalized with both acetate and picolinate pendant arms (Chart 1). The first member of this series to be investigated was H₄DO3APA, which forms lanthanide complexes with thermodynamic stabilities and kinetic inertness clearly inferior to those of DOTA.²⁴ More recently, we also reported the 1,4-H₃DO2APA and 1,7-H₃DO2APA ligands, whose lanthanide complexes were shown to possess some dramatically different properties in spite of having identical donor sets.²⁵ For instance, the exchange rate of the coordinated water molecule in [Gd(1,4-DO2APA)] was found to be three orders of magnitude faster than that of [Gd(1,7-DO2APA)], which was attributed to the weak coordination of the inner-sphere water molecule in the former due to the capping bond effect.^{25–27} In subsequent works we also showed that Y³⁺ and Gd³⁺ complexes of regioisomeric pyclen-based ligands showed rather different thermodynamic stabilities and dissociation kinetics.^{28,29}

Chart 1. Structures of the ligands discussed in this work.



Given the very different exchange kinetics of the coordinated water molecules observed for the Gd³⁺ complexes of 1,4-DO2APA and 1,7-DO2APA, we sought to analyze whether the different arrangement of the ligand donor atoms around the Ln³⁺ ions also had an effect on the thermodynamic stability and dissociation kinetics of the complexes. Thus, herein we report protonation constants of the ligands and the stability constants of the metal complexes formed with several Ln³⁺ and divalent ions available in vivo (Mg²⁺, Ca²⁺, Zn²⁺ and Cu²⁺). For this purpose, we employed a combination of potentiometric, relaxometric and spectrophotometric measurements. The dissociation kinetics of the Ln³⁺ complexes was investigated by following the exchange reactions of the correspond-

ing complexes with Cu²⁺ at different pH values. A computational study was performed to gain insight at the molecular level into the trends observed across the lanthanide series. We will show for the first time that the rates of dissociation can vary dramatically across the lanthanide series, and that the different arrangement of the donor atoms around the metal ion may result in dissociation rates differing up to six orders of magnitude. Finally, we report paraCEST studies that provided the exchange rate of the coordinated water molecule in [Yb(1,7-DO2APA)].

RESULTS AND DISCUSSION

Ligand protonation constants. The protonation constants of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻, defined as in eqn (1), were determined using pH-potentiometric titrations in 0.15 M NaCl. The logK_i^H values are compared with those of DO3APA⁴⁻ and DOTA⁴⁻ in Table 1.^{24,30}

$$K_i^H = \frac{[H_iL]}{[H_{i-1}L][H^+]} \quad \text{with } i = 1, 2, \dots, 5 \quad (1)$$

The protonation constants of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ are very similar, as expected for two ligands having identical nature of their protonation sites. The first and second protonation processes are associated to two nitrogen atoms of the macrocycle in *trans* position, which allows a larger charge separation of the protonated sites.³¹ The remaining three protonation constants are attributed to the stepwise protonation of the three carboxylate groups of the ligand or additional N atoms of the macrocyclic unit. The first protonation constant is two logK units lower than that reported for DOTA⁴⁻, which might be related, at least in part, to the higher negative charge of the latter. The total basicity ($\sum \log K_i^H$, $i = 1-4$) of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ is somewhat lower than that of DOTA⁴⁻, but slightly higher than for DO3APA³⁻ (Table 1). The values of the first protonation constants logK₁^H are significantly reduced on increasing the concentration of the background electrolyte from 0.15 M to 1 M (NaCl), which can be attributed to the formation of relatively stable Na⁺ complexes with these ligands.²⁴

Stability constants of the metal complexes. The stability constants of the Mg²⁺, Ca²⁺ and Zn²⁺ complexes with 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ were determined using direct pH-potentiometric titrations. The stability constants are defined as in eqn (2), while the protonation constants of the complexes are expressed in eqn (3):

$$K_{ML} = \frac{[ML]}{[M][L]} \quad (2)$$

$$K_{MH_iL} = \frac{[MH_iL]}{[MH_{i-1}L][H^+]} \quad (3)$$

The two ligands form rather weak complexes with Mg²⁺ and Ca²⁺, the corresponding stability constants being several orders of magnitude lower than those of DOTA⁴⁻ and DO3APA⁴⁻ (Table 1). This comparison should be taken with caution due to the different ionic strengths used for stability constant determination. However, the Zn²⁺ complexes present

thermodynamic stabilities comparable to those of the DOTA⁴⁻ and DO3APA⁴⁻ analogues. It is noteworthy that the Ca²⁺ complex of 1,4-DO2APA³⁻ is almost four orders of magnitude more stable than that of 1,7-DO2APA³⁻.

The stability constants of the Cu²⁺ complexes could not be determined by direct potentiometric titrations, as the formation of a protonated complex was quantitative at the beginning of the titrations at pH~1.7. Thus, the formation of the complexes was followed by monitoring the changes of the *d-d* absorption bands of the metal/ligand systems as a function of proton concentration. At pH~1 the two systems present an absorption band in the visible spectrum at ca. 705 nm ($\epsilon = 125 \text{ M}^{-1}\text{cm}^{-1}$)

whose intensity decreases upon increasing proton concentration (Fig. 1, see also Fig. S1, Supporting Information). Under these conditions, the absorption spectra are very similar to that measured for the Cu²⁺:DO3A³⁻ system,³² suggesting that the diprotonated form of the complex dominates the solution speciation at this pH possesses a similar structure. In the case of [Cu(H₂DOTA)],³³ an X-ray structure evidenced coordination of the four nitrogen atoms of the macrocycle and two oxygen atoms of carboxylate groups, while the two uncoordinated arms are protonated.

Table 1. Ligand protonation constants as well as stability and protonation constants of metal complexes of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ determined by pH-potentiometry (0.15 M NaCl, 25 °C) and literature data for related systems.

Metals / Ligand	1,4-DO2APA ³⁻	1,7-DO2APA ³⁻	DO3APA ^{4-,c}	DOTA ^{4-,d}	
	$\log K_1^H$	10.21(1); 9.44(5) ^e	10.15(2); 9.25(6) ^e	9.21	12.09
	$\log K_2^H$	9.46(1); 9.55(2) ^e	9.26(2); 9.32(2) ^e	8.94	9.76
	$\log K_3^H$	4.17(2); 4.19(5) ^e	4.22(2); 4.29(5) ^e	4.82	4.56
	$\log K_4^H$	3.65(3); 3.82(4) ^e	3.25(3); 3.49(4) ^e	3.52	4.09
	$\log K_5^H$	1.95(3); 1.95(3) ^e	2.38(3); 2.41(3) ^e	1.39	-
	$\sum \log K_i^H (i = 1-4)$	27.49; 27.00 ^e	26.88; 26.35 ^e	26.49	30.50
Mg	$\log K_{\text{MgL}}$	6.59(1)	7.25(8)	10.44	11.91
	$\log K_{\text{MgHL}}$	-	-	6.89	3.92
	$\log K_{\text{MgH2L}}$	-	-	6.37	-
	pM ^h	6.0	6.0	8.2	6.2
Ca	$\log K_{\text{CaL}}$	12.72(1)	8.96(7)	14.82	17.23
	$\log K_{\text{CaHL}}$	4.28(5)	-	4.59	3.54
	$\log K_{\text{CaH2L}}$	-	-	4.32	4.19
	pM ^h	8.8	6.1	12.4	11.1
Cu	$\log K_{\text{CuL}}$	23.84(3) ^{a,e}	23.52 ^{a,e}	23.20	22.25
	$\log K_{\text{CuHL}}$	4.10(2)	3.93(3)	4.17	3.78
	$\log K_{\text{CuH2L}}$	2.97(1)	2.80(3)	3.31	3.77
	$\log K_{\text{CuH3L}}$	-	-	1.97	-
	pM ^h	19.9	19.9	20.8	16.2
Zn	$\log K_{\text{ZnL}}$	19.36(4)	20.23(4)	20.25	21.09
	$\log K_{\text{ZnHL}}$	3.92(3)	3.55(3)	4.42	4.18
	$\log K_{\text{ZnH2L}}$	3.28(1)	2.84(1)	3.06	3.52
	$\log K_{\text{ZnH3L}}$	-	-	1.98	-
	pM ^h	15.4	16.6	17.8	15.0
La	$\log K_{\text{LaL}}$	15.96(1)	15.26(1)	21.17	21.7 ^f
	$\log K_{\text{LaHL}}$	-	-	2.55	-
	pM ^h	12.0	11.6	18.8	15.6
Gd	$\log K_{\text{GdL}}$	16.98(2), 17.12(2) ^b	16.33(2) ^a , 16.69(2) ^b	23.31	24.7 ^f , 25.3 ^g
	$\log K_{\text{GdHL}}$	-	-	2.65	-
	pM ^h	13.1	12.7	20.9	18.6
Lu	$\log K_{\text{LuL}}$	19.6(2), 19.21(2) ^b	15.24(2) ^a , 14.55(2) ^b	22.82	25.4 ^f
	$\log K_{\text{LuHL}}$	-	-	2.48	-
	pM ^h	15.7	11.6	20.4	19.3

^a Determined by simultaneous fitting of pH-potentiometric and UV-vis spectrophotometric data. ^b Determined by ¹H relaxometry. ^c Data in 0.1 M KCl from references 24 and 34. ^d Data in 0.1 M Me₄NNO₃ from references 30 and 35. ^e Data determined using a 1 M (C_{HCl}+C_{NaCl}) ionic strength. ^f Data in 0.1 M NaCl from reference 38. ^g Data from reference 39. ^h pM are defined as $-\log[M]_{\text{free}}$ for $[M]_{\text{tot}} = 1 \mu\text{M}$ and $[L]_{\text{tot}} = 10 \mu\text{M}$.

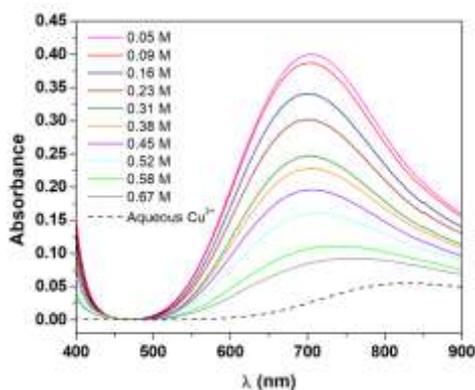


Figure 1. VIS spectra of the solutions containing nearly equimolar amounts of Cu^{2+} and 1,4-DO2APA ($c_{\text{Cu}^{2+}}=2.99$ mM, $c_{\text{Lig}}=3.03$ mM) as a function of acid concentration (see inset). The dashed spectrum corresponds to aqueous Cu^{2+} in 1.0 M NaCl.

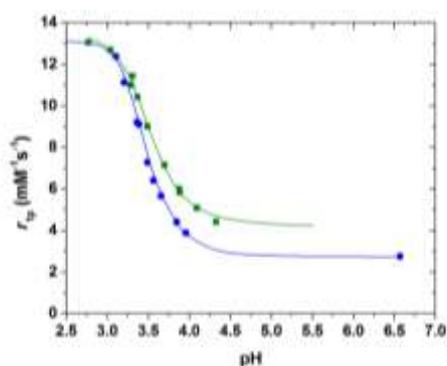


Figure 2. Relaxivity pH profiles of the equilibrated $[\text{Gd}(1,4\text{-DO2APA})]$ (squares) and $[\text{Gd}(1,7\text{-DO2APA})]$ (circles) complexes ($[\text{Gd}(\text{L})] = 1.5 \times 10^{-3}$ M normalized to 1.0×10^{-3} M, $T = 25$ °C, $I = 0.15$ M NaCl, $B = 0.47$ T). The solid lines correspond to the fits of the data to obtain the stability constants given in Table 2.

The simultaneous analysis of the spectroscopic and potentiometric data afforded the stability and protonation constants reported in Table 1. The $\log K_{\text{CuL}}$ values obtained for the complexes of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ are slightly higher than those reported for DO3APA⁴⁻³⁴ and DOTA^{4-30,35}. The corresponding pM values of the complexes containing picolinate groups are clearly higher than that calculated for the complex with DOTA⁴⁻. The speciation diagrams (Fig. S2-S8, Supporting Information) evidence the formation of protonated complex species below pH ~ 6 . Complex dissociation occurs only at rather acidic pH values ($< \text{pH} \sim 3$).

The stability constants of representative Ln^{3+} complexes were obtained by using both pH-potentiometry and the relaxometric technique (Gd^{3+}),³¹ or by potentiometric titrations only (La^{3+} and Lu^{3+}). The relaxivity of the $[\text{Gd}(1,4\text{-DO2APA})]$ complex measured at 20 MHz and 25 °C ($\text{pH} = 4.5$) amounts to $r_{1p} = 4.2$ $\text{mM}^{-1}\text{s}^{-1}$, in good agreement with that reported previously.²⁵ In the case of $[\text{Gd}(1,7\text{-DO2APA})]$ the relaxivity measured in the same conditions ($r_{1p} = 2.7$ $\text{mM}^{-1}\text{s}^{-1}$) is somewhat higher than that measured previously without controlling the ionic strength ($r_{1p} = 2.0$ $\text{mM}^{-1}\text{s}^{-1}$). This is likely due to a faster water exchange of the coordinated water molecule in 0.15 M NaCl, which is expected to increase the inner-sphere contribution to relaxivity.³⁶ The

relaxivities of the two complexes increase below pH 4.5 reaching a value of 13.0 $\text{mM}^{-1}\text{s}^{-1}$ (Fig. 2) at pH 3, which corresponds to the relaxivity of $[\text{Gd}(\text{H}_2\text{O})_8]^{3+}$.³⁷ The analysis of the relaxivity data provided the stability constants of the two Gd^{3+} complexes, which turned out to be rather similar (Table 1). The stability constants determined using relaxometry and pH-potentiometry are in good agreement. The stability constants determined for the Gd^{3+} complexes resulted to be six-seven orders of magnitude lower than that reported for $[\text{Gd}(\text{DO3APA})]^{-}$,²⁴ and seven-eight orders of magnitude lower than the stability constants reported for $[\text{Gd}(\text{DOTA})]^{-}$.^{38,39} This highlights the dramatic effect that removing one of the carboxylate groups of the ligand causes on the thermodynamic stability of the complexes. The stability constants of the complexes with 1,4-DO2APA³⁻ increase across the lanthanide series from La^{3+} to Lu^{3+} , with $\Delta \log K = \log K_{\text{LuL}} - \log K_{\text{LaL}} = 3.2$. This trend is observed for most lanthanide complexes as a consequence of the increasing charge density of the metal ion due to the lanthanide contraction.⁴⁰

The stability constants of 1,4-DO2APA³⁻ complexes increase along the series, as usually observed for Ln^{3+} complexes. In contrast, the stability of the complexes with 1,7-DO2APA³⁻ increases one order of magnitude from La^{3+} to Gd^{3+} and then decreases from Gd^{3+} to Lu^{3+} . This stability trend is rather uncommon, though has been observed for the complexes of DTPA⁵⁻⁴¹. A fully reversed stability trend with the stability decreasing across the series was also observed for a few systems, and attributed to the weakening of some Ln^{3+} -donor bonds due to steric constraints.⁴² As a result, a smaller increase of the overall electrostatic interaction between the ligand and the Ln^{3+} ion across the 4f period does not compensate the more negative hydration energies of the lanthanide ions.⁴⁰

Dissociation kinetics. Both the free metal ion and the ligand released as a result of complex dissociation are potentially toxic (the former being more toxic), and thus metal-complexes for medical applications must have high kinetic inertness. The inertness of metal complexes with respect to dissociation is often assessed by measuring dissociation rates under acidic conditions, sometimes in the presence of Zn^{2+} or Cu^{2+} metal ions as ligand scavengers. Generally, the dissociation of Gd^{3+} complexes with cyclen-based ligands such as DOTA⁴⁻ or DO3A³⁻ occurs slowly following the proton-assisted dissociation mechanism, while the presence of metal ions such as Zn^{2+} and Cu^{2+} in most of the cases does not affect the dissociation rates.^{43,44}

The inertness of the $[\text{Ln}(1,4\text{-DO2APA})]$ and $[\text{Ln}(1,7\text{-DO2APA})]$ complexes ($\text{Ln} = \text{La}, \text{Gd}$ or Yb) was assessed by studying the rates of the metal exchange reactions occurring with Cu^{2+} ions (Eq 4) at 25 °C by following the formation of the corresponding Cu^{2+} complex at 300 nm (Eq (4)).



Where L denotes either the 1,4-DO2APA³⁻ or 1,7-DO2APA³⁻ chelator. In the presence of 10-40 fold excess of the Cu^{2+} -ion, the transmetalation reaction of the complexes can be treated as a pseudo-first-order process whose rate can be expressed by eqn (5):

$$-\frac{d[\text{LnL}]_t}{dt} = k_d[\text{LnL}]_t \quad (5)$$

where k_d is a pseudo-first-order rate constant and $[\text{LnL}]_t$ is the concentration of the LnL species at time t .

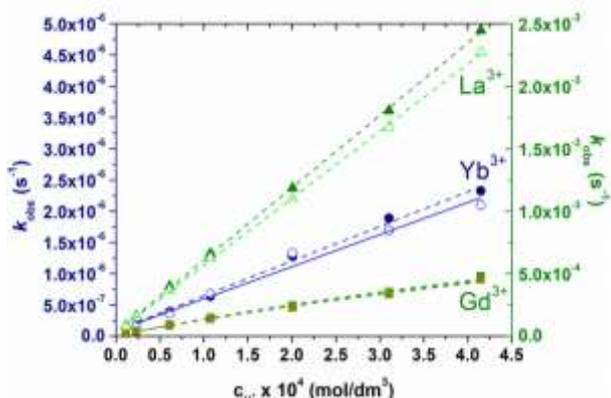


Figure 3. Pseudo-first-order rate constants of dissociation of $[\text{Ln}(1,4\text{-DO2APA})]$ complexes ($\text{Ln}^{3+} = \text{La}^{3+}$ (triangles, shades of green), Gd^{3+} (squares, shades of green) or Yb^{3+} (circles, shades of blue)) as a function of H^+ ion concentration measured by using a large excess of exchanging Cu^{2+} ion (for Gd $c_{\text{Cu}^{2+}} = 3.33, 5.95, 9.28,$ and 11.9 mM corresponding to 11.5, 20.5, 30 and 41 fold metal excess, respectively; in other Ln^{3+} systems 11 (3.37 mM, closed symbols) and 31 (10.1 mM, open symbols) fold Cu^{2+} excess was applied).

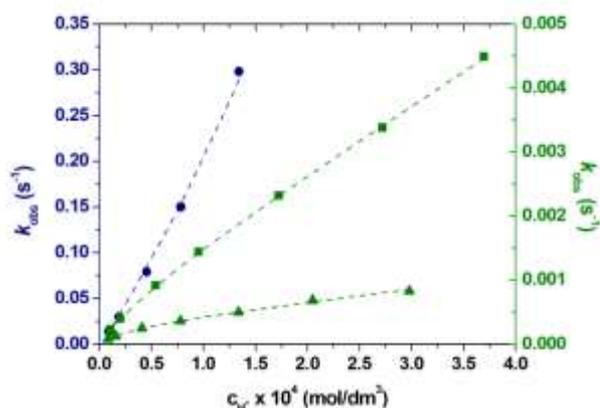
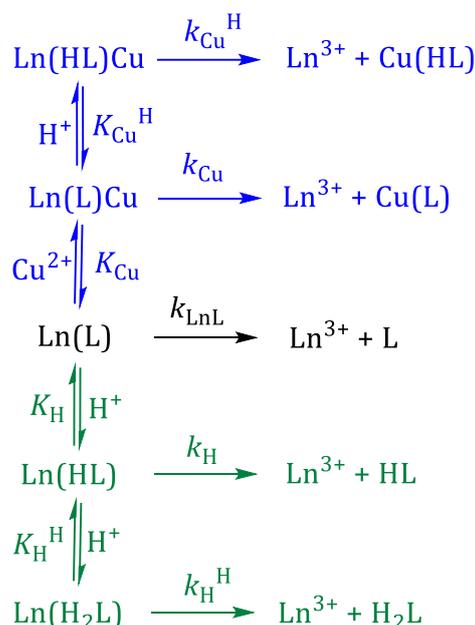


Figure 4. Pseudo-first-order rate constants of dissociation of $[\text{Ln}(1,7\text{-DO2APA})]$ complexes ($\text{Ln}^{3+} = \text{La}^{3+}$ (green triangles), Gd^{3+} (green squares) or Yb^{3+} (blue circles)) as a function of H^+ ion concentration measured by using large excess of exchanging Cu^{2+} ion ($c_{\text{Cu}^{2+}} = 2.02$ mM; 10.1 fold metal excess). The lines represent the fits of the data as explained in the text.

The rates of the transmetallation reactions of $[\text{Ln}(1,4\text{-DO2APA})]$ complexes were found to be either directly proportional to the H^+ concentration ($\text{Ln}=\text{La}^{3+}$ and Yb^{3+}) or showed saturation-like behavior (Gd^{3+}). For all systems studied an increase in Cu^{2+} concentration was found to slow the dissociation (Figure 3, see also Figure S9, Supporting Information). For the $[\text{Ln}(1,7\text{-DO2APA})]$ complexes ($\text{Ln}=\text{La}^{3+}$ and Gd^{3+}) saturation-like behavior was observed with hydrogen ion concentration, while the Yb^{3+} system shows a second order dependence on hydrogen ion concentration (Fig. 4). The pseudo-first-order rate constants characterizing the dissociation of the Ln^{3+} complexes formed with the 1,7-DO2APA ligand are independent of the Cu^{2+}

ion concentration and thus kinetic experiments were performed while keeping the Cu^{2+} ion concentration constant (20 fold excess was utilized in all kinetic experiments). The total concentration of the complexes can be expressed by considering the species present in solution under the conditions applied in the kinetic studies. Obviously the majority of the complex is present in its LnL form, but lowering the pH may result in the formation of mono- and diprotonated intermediate species (LnHL and LnH_2L), while in the presence of a scavenging metal ion (such as Cu^{2+} used in the studies), the formation of a dinuclear intermediate (LnLCu) may occur by involving weakly bound (to Ln^{3+} ions) or uncoordinated donor atom(s) in Ln^{3+} complexes (the latter complex may even protonate to form a $\text{Ln}(\text{HL})\text{Cu}$ intermediate). These reaction pathways are summarized in Chart 2.

Chart 2. Dissociation mechanisms of LnL complexes.



The total concentration of the LnL complexes in kinetic experiments ($[\text{LnL}]_{\text{tot}}$) can be expressed as the sum of the concentrations of all the different Ln-containing species present in solution:

$$[\text{Ln}(\text{L})]_{\text{tot}} = [\text{LnL}] + [\text{Ln}(\text{HL})] + [\text{Ln}(\text{H}_2\text{L})] + [\text{Ln}(\text{L})\text{Cu}] + [\text{Ln}(\text{HL})\text{Cu}] \quad (6)$$

Taking this into account and using eqn (5) the following rate equation is obtained:

$$-\frac{d[\text{Ln}(\text{L})]_{\text{tot}}}{dt} = k_d[\text{Ln}(\text{L})]_{\text{tot}} = k_{\text{LnL}}[\text{Ln}(\text{L})] + k_{\text{H}}[\text{Ln}(\text{HL})] + k_{\text{H}}^{\text{H}}[\text{Ln}(\text{H}_2\text{L})] + k_{\text{Cu}}[\text{Ln}(\text{L})\text{Cu}] + k_{\text{Cu}}^{\text{H}}[\text{Ln}(\text{HL})\text{Cu}] \quad (7)$$

The rate constants in eqn (7) characterize the rate of the spontaneous (k_{LnL}), proton-assisted ($k_{\text{H}}, k_{\text{H}}^{\text{H}}$), copper-assisted (k_{Cu}) and proton-copper-assisted (k_{Cu}^{H}) dissociation pathways. Taking into account the different reaction pathways shown in Chart 2 and the equations determining the $K_{\text{H}}, K_{\text{H}}^{\text{H}}, K_{\text{Cu}}$ and K_{Cu}^{H} equilibrium constants, the pseudo-first-order rate constant (k_d) can be expressed by eqn (8):

$$k_d = \frac{k_{\text{LnL}} + k_1[\text{H}^+] + k_2[\text{H}^+]^2 + k_3[\text{Cu}^{2+}] + k_4[\text{Cu}^{2+}][\text{H}^+]}{1 + K_{\text{H}}[\text{H}^+] + K_{\text{H}}^{\text{H}}[\text{H}^+]^2 + K_{\text{Cu}}[\text{Cu}^{2+}] + K_{\text{Cu}}^{\text{H}}[\text{Cu}^{2+}][\text{H}^+]} \quad (8)$$

where $K_H = [\text{Ln}(\text{HL})]/[\text{Ln}(\text{L})][\text{H}^+]$, $K_{\text{Cu}} = [\text{Ln}(\text{L})\text{Cu}]/[\text{Ln}(\text{L})][\text{Cu}]$, $K_{\text{Cu}}^{\text{H}} = [\text{Ln}(\text{HL})\text{Cu}]/[\text{Ln}(\text{L})\text{Cu}][\text{H}^+]$, $k_1 = k_{\text{H}} \cdot K_H$, $k_2 = k_{\text{H}}^{\text{H}} \cdot K_{\text{H}} \cdot K_{\text{H}}^{\text{H}}$, $k_3 = k_{\text{Cu}} \cdot K_{\text{Cu}}$, and $k_4 = k_{\text{Cu}}^{\text{H}} \cdot K_{\text{Cu}}^{\text{H}}$. Eqn (8) takes into account all the rational/possible dissociation pathways expected to occur under the applied conditions. As for other systems, this expression can be simplified to fit the rate constants determined as a function of pH (for both systems) and Cu^{2+} ion concentration (for the complexes of dissociation of 1,4-DO2APA). For instance, the rate of the spontaneous ($k_{\text{L,ML}}$) dissociation of complexes with macrocyclic ligands derived from DOTA⁴⁻ is generally extremely low, and therefore its accurate determination is very challenging (for the $[\text{Gd}(\text{DOTA})]^-$ the values lie in the range of 5×10^{-8} to $6.7 \times 10^{-11} \text{ s}^{-1}$).^{19,45} In the present work, we could obtain reliable data for the rate of spontaneous dissociation only in a few cases (Table 2). For all other systems the given rate constant was fixed to 0 during the data refinement. On the other hand, involvement of the metal ion in the dissociation of 1,7-DO2APA complexes could not be confirmed and thus the terms characterizing this pathway (rate constants k_{Cu} and k_{Cu}^{H} and equilibrium constants K_{Cu} and K_{Cu}^{H}) were neglected during the data refinement. For the complexes of 1,4-DO2APA³⁻ the value of the pseudo-first-order rate constant decreases with increasing Cu^{2+} concentration, indicating that the dinuclear intermediate formed in the reaction acts as a “dead-end” complex. In agreement with this phenomenon the data fitting returned the stability constants of the dinuclear intermediates (K_{Cu}), but the rate constants of the metal-assisted dissociation had to be neglected. Altogether, these observations agree with literature data, where the acid assisted dissociation (characterized by k_1 and k_2 rate constants) was recognized to be the most important pathway for the dissociation of Ln^{3+} complexes formed with DOTA⁴⁻ and its derivatives.

The rate constants obtained by fitting the k_d first-order dissociation rate constants are presented and compared in Table 2. The results show that the complexes of the two regioisomeric ligands behave very differently as far as their dissociation kinetics concerns. The comparison of the k_1 rate constants shows that removal of an acetate pendant arm from the DO3APA⁴⁻ ligand results in a considerable drop in the inertness of the complexes formed with Ln^{3+} ions.²⁴ Indeed, the Gd^{3+} complexes of DO2APA³⁻ exhibit rate constants 3-4 orders of magnitude higher than that of the $\text{Eu}(\text{DO3APA})$ chelate. This is not unexpected, since a similar increase in the rate constants characterizing the acid-assisted dissociation process can be observed when the data of $\text{Gd}(\text{DOTA})$ and $\text{Gd}(\text{DO3A})$ are compared.³² More striking are the differences observed when the data of the two isomeric complexes are compared. The complexes of the 1,4-regioisomer behave very much like the majority of the complexes formed with DOTA⁴⁻ and DOTA derivative ligands: the rate constants of acid-assisted dissociation decrease with decreasing size of the Ln^{3+} ions. However, the complexes of the 1,7-regioisomer become considerably less inert as the size of the Ln^{3+} ions gets smaller due to the lanthanide contraction. As an interplay between the two opposite trends, the ratios of the k_1 values of the complexes with the two ligands are 2, 16 and 2.5×10^5 for the La^{3+} , Gd^{3+} and Yb^{3+} complexes, respectively. To the best of our knowledge, the complexes of 1,7-DO2APA³⁻ represent the first example of Ln^{3+} complexes with DOTA derivatives that show an increase in the rate constants of acid-assisted dissociation across the lanthanide series. The different trends observed for the complexes of 1,4-DO2APA and 1,7-DO2APA are very likely related to the labile capping bond effect, which was introduced to explain the differences in terms of water exchange rates of Gd^{3+} complexes formed with the 1,4- and 1,7-regioisomers investigated here.²⁵

Table 2. Dissociation rate and equilibrium constants characterizing the dissociation of the $[\text{Ln}(1,4 \text{ DO2APA})]$ and $[\text{Ln}(1,7 \text{ DO2APA})]$ complexes ($I = 0.15 \text{ M NaCl}$, $25 \text{ }^\circ\text{C}$, $\text{Ln} = \text{La}$, Gd or Yb) and related systems provided for comparison.

Metal (M) / Ligand	1,4-DO2APA ^{3-a}	1,7-DO2APA ^{3-a}	DO3APA ^{4-b}	DOTA ^{4-c}	
La^{3+}	$k_0, \text{ s}^{-1}$	0.15 M NaCl	0.15 M NaCl	0.1 M KCl	0.15 M NaCl
	$k_1, \text{ M}^{-1}\text{s}^{-1}$	$3.1(7) \times 10^{-5}$	–	–	–
	$k_2, \text{ M}^{-2}\text{s}^{-1}$	6.02(5)	12.2(7)	$2.40 \times 10^{-3} (\text{Ce}^{3+})$	$8.0 \times 10^{-4} (\text{Ce}^{3+})$
	K_{H}	–	$7(2) \times 10^5$	–	$2.0 \times 10^{-3} (\text{Ce}^{3+})$
	K_{Cu}	–	$3.5(7) \times 10^4$	–	–
Gd^{3+}	$k_0, \text{ s}^{-1}$	12(1)	–	–	–
	$k_0, \text{ s}^{-1}$	$5.8(6) \times 10^{-6}$	–	–	$6.7 \times 10^{-11};^d$ $< 5 \times 10^{-8};^e$ $5.8 \times 10^{-10} \quad (37$ $^\circ\text{C});^f$
	$k_1, \text{ M}^{-1}\text{s}^{-1}$	1.42(4)	23.5(5)	$1.56 \times 10^{-3} (\text{Eu}^{3+})$	$1.8 \times 10^{-6};^d$ $8.4 \times 10^{-6};^e$ $2 \times 10^{-5} (37 \text{ }^\circ\text{C});^f$
	$k_2, \text{ M}^{-2}\text{s}^{-1}$	–	$2.1(3) \times 10^5$	$4.8 \times 10^{-4} (\text{Eu}^{3+})$	–
	K_{H}	$7.9(9) \times 10^2$	$2.0(3) \times 10^4$	–	–
Yb^{3+}	K_{Cu}	6(2)	–	–	–
	$k_0, \text{ s}^{-1}$	$1.0(5) \times 10^{-7}$	–	–	–
	$k_1, \text{ M}^{-1}\text{s}^{-1}$	$5.8(3) \times 10^{-3}$	$1.48(7) \times 10^3$	2.79×10^{-3}	–
	$k_2, \text{ M}^{-2}\text{s}^{-1}$	–	$6(1) \times 10^6$	–	–
	K_{H}	–	–	–	–
K_{Cu}	13(8)	–	–	–	

^a This work. Values within parenthesis represent the statistical errors of the fits in the last significant digits. ^b Data from Ref 24. ^c Data from Ref 46. ^d Data from Ref 32. ^e Data from Ref 45. ^f Data from Ref 43.

Complexes with regioisomeric ligands derived from a pyclen platform were also found to present significant differences in terms of kinetic inertness.^{28,29} However, the amplitude of the differences in the dissociation rates are considerably less pronounced for the pyclen derivatives. Furthermore, pyclen derivatives present k_1 values four orders of magnitude lower. This can be rationalized in terms of different rigidity of the parent macrocycles (cyclen vs. pyclen). The pyclen macrocycle is considerably more rigid than cyclen, and thus the energy cost required to rearrange the complex structure during the dissociation process is expected to be higher for the more rigid pyclen derivatives.

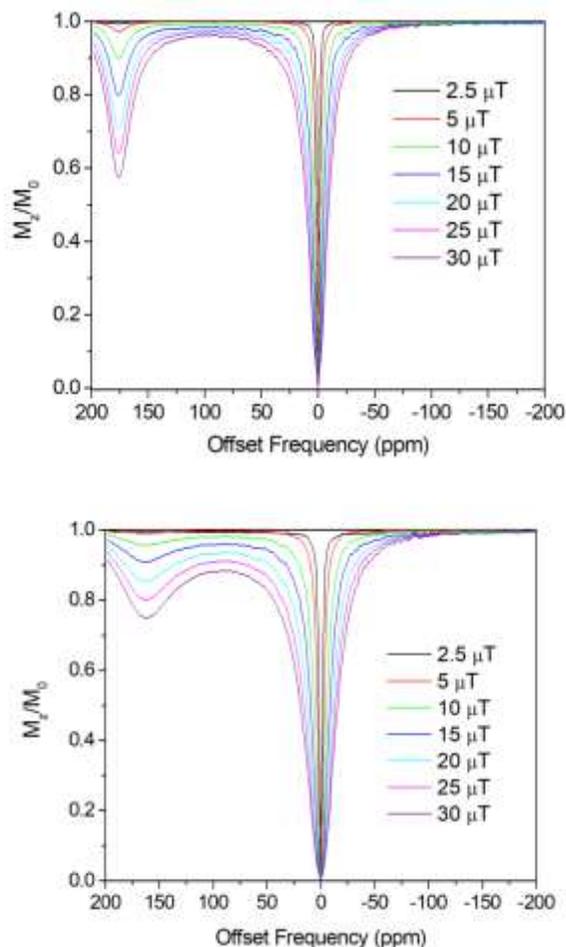


Figure 5. CEST spectra of [Yb(1,7-DO2APA)] in H₂O (15 mM, pH 7, saturation time 10 s) recorded using different saturation powers at 25 (top) and 37 °C (bottom).

The Z-spectra (25 °C, saturation time 10 s) recorded from a 15 mM solution of [Yb(1,7-DO2APA)] present a prominent CEST feature at 176 ppm that is already visible at low saturation powers (3% at $B_1 = 5 \mu\text{T}$) and reaches 33% at $B_1 = 30 \mu\text{T}$ (Fig. 5). The CEST peak becomes broader and shifts to 162 ppm at 37 °C (29% at $B_1 = 30 \mu\text{T}$, Fig. S10, Supporting Information). The analysis of the CEST spectra using Bloch equations⁵³ provides a k_{ex}^{298} value of $(15.9 \pm 1.6) \times 10^3 \text{ s}^{-1}$. This

value is similar to that estimated for the Gd³⁺ analogue using ¹H NMRD and ¹⁷O NMR measurements ($k_{\text{ex}}^{298} = 8.8 \times 10^3 \text{ s}^{-1}$).¹² Water exchange is faster at 37 °C with a k_{ex}^{310} value of $(51.6 \pm 0.9) \times 10^3 \text{ s}^{-1}$. Thus, CEST experiments confirm the very slow water exchange of the coordinated water molecule in [Ln(1,7-DO2APA)] complexes, as a result of a strong Ln-O_{water} interaction. Slow water exchange rates were also reported recently for [Ln(AAZTA)(H₂O)]⁻ (Ln = Tm, Yb) derivatives in spite of their negative charge.^{49,54} These results represent another evidence that positively charged Ln³⁺ complexes (or their derivatives) are not a prerequisite to achieve slow water exchange rates, useful for the design of paraCEST agents.

CEST studies. Chemical exchange saturation transfer experiments allow estimating the exchange rates of protons in slow-to-intermediate exchange with bulk water.⁴⁷⁻⁴⁹ Among the different Ln³⁺ ions, Yb³⁺ induces relatively large pseudocontact shifts without extensive line broadening due to paramagnetic relaxation effects.⁵⁰⁻⁵² In a previous work, we showed that the coordinated water molecule present in [Gd(1,7-DO2APA)] shows a very low water exchange rate.²⁵ Thus, we performed CEST experiments using the Yb³⁺ analogue to assess the water exchange rate of the coordinated water molecule and provide further experimental evidence of the labile capping bond phenomenon.

Computational study. The unusual variation of the stability constants and dissociation rate constants across the lanthanide series prompted us to perform a computational study to gain insight into the structural origin of these effects at the molecular level. The structures of the [Ln(1,4-DO2APA)(H₂O)]·2H₂O and [Ln(1,7-DO2APA)(H₂O)]·2H₂O systems (Ln = La, Gd or Yb) were optimized at the TPSSH/LCRECP/6-31G(d,p) level following the methodology described before⁵⁵ (see also computational details below). The model systems investigated include two second-sphere water molecules, which were found to be crucial to provide a good description of the environment around the coordinated water molecule.⁵⁶ These water molecules are involved in hydrogen-bonding interactions with the inner-sphere water molecule and oxygen atoms of carboxylate groups, but they do not coordinate to the metal ion. Our previous relaxometric and luminescence study pointed to the presence of a hydration equilibrium in [Gd(1,4-DO2APA)] involving a nine-coordinated species with an inner-sphere water molecule and an eight-coordinated species lacking the coordinated water molecule.²⁵ We therefore initiated the computational study by exploring the potential energy surface of the [Ln(1,4-DO2APA)(H₂O)]·2H₂O systems (Ln = La, Gd or Yb) by increasing the distance between the metal ion and the oxygen atom of the coordinated water molecule. The corresponding relaxed potential energy surface scans are shown in Fig. 6. Our results indicate that the hydrated species is clearly more stable for the La³⁺ complex, while the situation is reversed for Yb³⁺. In the case of the Gd³⁺ complex the $q = 1$ and $q = 0$ species (q is the number of coordinated water molecules) present very similar energies, which is in nice agreement with

the hydration number of $q = 0.55$ estimated from the analysis of ^1H NMRD and ^{17}O NMR data.²⁵

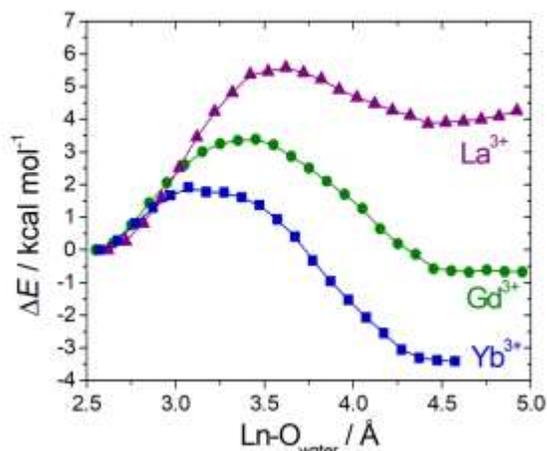


Figure 6. Relaxed potential energy surface scans computed for the $[\text{Ln}(1,4\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ systems ($\text{Ln} = \text{La}, \text{Gd}$ or Yb) by increasing the distance between the lanthanide ion and the oxygen atom of the coordinated water molecule in steps of 0.1 \AA .

Fig. 7 presents the optimized geometries of the $[\text{Yb}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ and $[\text{Yb}(1,4\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ complexes and the numbering scheme adopted for the donor atoms of the ligands. The $[\text{Yb}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ complex presents a capped square antiprismatic coordination geometry in which an oxygen atom of the picolinate group (O3) occupies the capping position. The steric compression around the capping position results in a rather long Yb-O3 distance (2.385 \AA) compared to the distances involving the oxygen atoms of acetate groups ($\text{Yb-O1} = 2.278 \text{ \AA}$; $\text{Yb-O2} = 2.320 \text{ \AA}$). This is in line with the capping bond effect introduced recently by our group.^{25,28,29} In $[\text{Yb}(1,4\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$, the capping position is occupied by the coordinated water molecule, which presents a very long Yb-O1w distance (2.483 \AA) compared to that obtained for the 1,7-DO2APA derivative (2.390 \AA). As a result, the most stable form of the Yb^{3+} complex with 1,4-DO2APA does not contain a coordinated water molecule. Thus, the change in hydration state on going from the La^{3+} complex ($q=1$) to Yb^{3+} ($q=0$) is a direct consequence of the steric compression around the capping water binding site. In $[\text{Ln}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ complexes the coordinated water molecule occupies a coordination position in the upper square plane of the SAP coordination polyhedron, and thus remains tightly bound throughout the lanthanide series from La^{3+} to Yb^{3+} .

The bond distances of the metal coordination environment in lanthanide complexes generally decrease across the lanthanide series due to the lanthanide contraction.⁵⁷ As a consequence, an analysis of the strength of the Ln-donor bonds in terms of bond distances is not straightforward. Thus, we calculated the

electron densities at the critical points of the Ln-donor bonds (ρ_{BCP}), which have been shown to provide a measure of the strength of the concerned bonds.^{26,27} The results of this analysis are presented in Fig. 8 (see also Table S1, Supporting Information). The Ln-O bonds involving carboxylate groups present higher ρ_{BCP} values than the Ln-N bonds, as would be expected considering the hard nature of the Ln^{3+} ions in Pearson's classification. All Ln-donor bonds are characterised by positive values of the Laplacian of the electron density at the bond critical points ($\nabla^2\rho_{\text{BCP}}$, Table S1, Supporting Information) and ρ_{BCP} values well below 0.2 au . This indicates that the electron density is locally depleted at the bond critical points, pointing to ionic interactions.^{58,59} The electron densities of the critical points at the Ln-O_{water} bonds are higher for the complexes of 1,7-DO2APA³⁻, which is in agreement with the slow water exchange rate measured for the Gd^{3+} complex.^{26,27}

For the $[\text{Ln}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ complexes, the ρ_{BCP} characterizing the donor atoms delineating the upper plane of the square antiprism (O1, O2 and O1w) increase across the lanthanide series, which indicates that the concerned bonds become stronger as a result of the increased positive charge density of the metal ion. However, the Ln-O3 bond, which involves the oxygen atom of the picolinate group, weakens across the lanthanide series due to an increasing compression around the capping bond. Thus, the weaker interaction of the metal ion with the picolinate oxygen atom appears to be responsible for the decrease in complex stability for the smaller Ln^{3+} ions. The same effect is likely responsible for the increasing acid-catalyzed dissociation rates observed across the 4f period. The acid-catalyzed dissociation of the complex proceeds through the formation of a protonated species involving negatively charged carboxylate groups of the ligands. Once protonated, a cascade of proton transfer events causes the protonation of the more basic amine nitrogen atoms, provoking complex dissociation. The relatively weak coordination of the picolinate group likely facilitates its decoordination upon protonation, so that the proton can be more easily transferred to the amine nitrogen atoms.

The trends obtained for the ρ_{BCP} values in the complexes of 1,4-DO2APA³⁻ indicate that all Ln-O bonds involving the carboxylate groups of the ligand become stronger across the lanthanide series, explaining the increasing stability of the complexes as well as the increasing kinetic inertness with respect to complex dissociation. The astonishing increase of the kinetic inertness on going from Gd^{3+} to Yb^{3+} is the result of the more compact structure of the complex associated to the reduced coordination number in the $q = 0$ Yb^{3+} complex. The coordinated water molecule in the complexes with 1,7-DO2APA becomes more tightly bound across the series, while the reverse situation is observed for the analogues of 1,4-DO2APA.

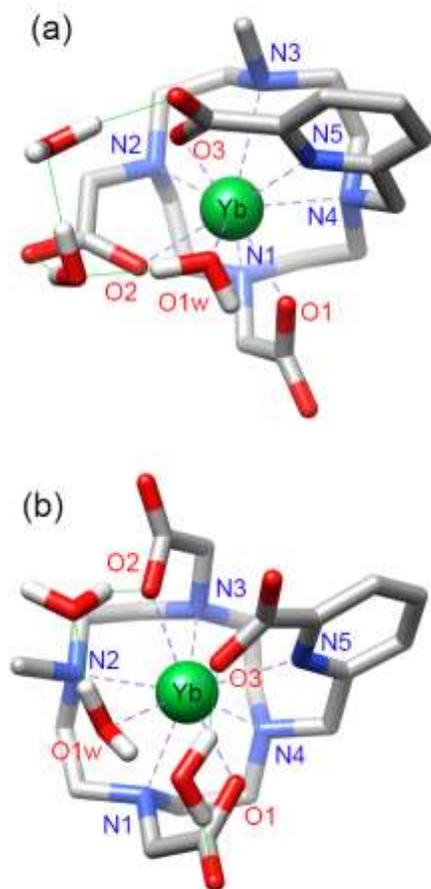


Figure 7. Optimized geometries of the $[\text{Yb}(1,4\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ (a) and $[\text{Yb}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ (b) systems obtained with DFT calculations.

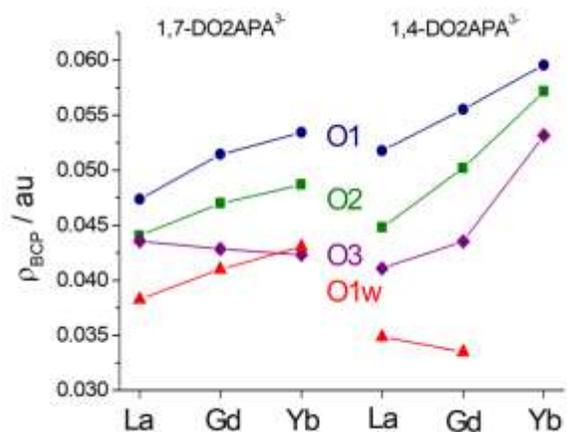


Figure 8. Electron densities at the bond critical points (ρ_{BCP}) involving Ln-O bonds in $[\text{Ln}(1,4\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ (Ln = La, Gd), $[\text{Yb}(1,4\text{-DO2APA})]$ and $[\text{Ln}(1,7\text{-DO2APA})(\text{H}_2\text{O})]\cdot 2\text{H}_2\text{O}$ (Ln = La, Gd, Yb) complexes.

CONCLUSIONS

We have presented a detailed analysis of the thermodynamic stability and dissociation kinetics of lanthanide complexes with two cyclen-based regioisomeric ligands. The two series of complexes present different stability trends across the lanthanide series: The complexes of 1,4-DO2APA show the typical increase in stability across the series, while the stability of the 1,7-DO2APA analogues decreases on going from Gd^{3+} to Lu^{3+} . The rates of dissociation of 1,4-DO2APA complexes decrease slightly from La^{3+} to Gd^{3+} and then drop dramatically for the Yb^{3+} complex, an effect that is related to a more tight coordination of the ligand, associated to the higher positive charge density of the metal ion and a reduced coordination number.

The rates of dissociation of the complexes with 1,7-DO2APA show an unprecedented increase on moving to the right along the lanthanide series. DFT studies show that this is likely related to the coordination of the picolinate oxygen atom at the sterically demanding apical position. The coordination of the picolinate group becomes weaker on decreasing the ionic radius of the Ln^{3+} ion, facilitating the decoordination of this group upon protonation. The coordination of the picolinate oxygen atom at the apical position leaves the water molecule tightly coordinated at the upper square plane of the square antiprismatic coordination polyhedron. As a result, the coordinated water molecule exhibits a low exchange rate with bulk water, as demonstrated by CEST experiments for the Yb^{3+} complex. While the low kinetic inertness of this complex precludes any potential application as MRI contrast agent, this result shows that the slow water exchange rates required for CEST agents may be attained placing water molecules at non-capping positions.

In summary, the arrangement of the donor atoms of the ligand around the Ln^{3+} ion plays a crucial role in the thermodynamic and kinetic properties of the complexes. We have shown that dissociation rates may vary up to three orders of magnitude across the lanthanide series. This unprecedented observation shows that the kinetic inertness of a complex may not be assured when one lanthanide ion is changed by another. Furthermore, the rates of dissociation of lanthanide complexes may vary by up to six orders of magnitude by changing the arrangement of the donor atoms around the metal ion. These issues must be carefully considered in ligand design in addition to the number and type of donor atoms.

EXPERIMENTAL AND COMPUTATIONAL SECTION

Equilibrium Measurements. The metal salts and other materials used in equilibrium experiments were of the highest analytical grade. Standardized $\text{Na}_2\text{H}_2\text{EDTA}$ was used to determine the concentration of the MgCl_2 , CaCl_2 , CuCl_2 , ZnCl_2 and LnCl_3 solutions by complexometric titration in the presence of eriochrome black T (MgCl_2), Patton & Reeder (CaCl_2), murexid (CuCl_2) and xylenol orange (ZnCl_2 and LnCl_3) as the indicators.

The concentration of the ligands was determined by pH-potentiometric titration in the presence and absence of a large excess of CaCl₂. To obtain the protonation constants of 1,4-DO2APA³⁻ and 1,7-DO2APA³⁻ acidified (HCl) ligand solutions (2 mM) were titrated with 0.15 M NaOH in the presence of 0.15 M NaCl as ionic strength. The stability and protonation constants of the Mg²⁺, Ca²⁺, Cu²⁺ (by using 1.0 M ionic background for this particular system) and Zn²⁺ complexes were determined by pH potentiometric titration where the metal to ligand ratios were 1:1 using 1.7 mM ligand concentration (100-200 V(mL) NaOH - pH data pairs were recorded). The pH-potentiometric titrations were carried out with a Metrohm 888 Titrando workstation with the use of a Metrohm 6.0234.100 combined electrode in the pH range of 1.70-11.80. For the calibration of the pH meter, KH-phtalate (pH = 4.005) and borax (pH = 9.177) buffers were used. The titrated samples (6 mL) were thermostated at 25 °C and mechanically stirred under inert atmosphere (N₂). The method proposed by Irving et al.⁶⁰ was used to calculate the H⁺ concentrations from the measured pH values. The Irving factor and the ion product of water (pK_w=13.75) was determined in the same HCl – NaOH titration experiment. In the case of the Cu²⁺ complexes pH-potentiometry was supported by spectrophotometric measurements performed on batch samples, since the formation of protonated Cu²⁺ complexes was found to be quantitative even near pH=1.5 (thus only the deprotonation events of the complexes could be characterized in the pH range of 1.70-11.80). The absorbance spectra were recorded with a Cary 100 Bio spectrophotometer at 25 °C, using 1.0 cm cells. For determining the stability constants of the Cu²⁺ complexes 8 separated samples were prepared with different acid concentrations (in the acid concentration range of 0.01 – 0.70 M) and the absorbance data measured at 7 wavelengths in the 650 - 800 nm range were used for the calculations. In these experiments, the ionic strength of the samples was set to 1.0 M (cHCl+cNaCl=1.0 M), and thus the protonation constants of the ligands were also determined under the same conditions (Table 1). Beside the pH-potentiometric data, the measured absorbance and molar absorptivity data determined independently were taken into account to determine the equilibrium constants.

Due to the relatively slow formation of the Ln³⁺ complexes, the so called "out-of-cell" technique was utilized to determine their stability constants. Samples with total volume of 2.50 mL containing the Ln³⁺ ion (2 mM) and the ligand (2 mM) were prepared in the pH range of about 3.0 – 6.5. The ionic strength in the samples was set to 0.15 M by using NaCl. The closed samples were kept at 25 °C for 4 weeks to reach the equilibrium and then their pH and relaxivity (for the Gd³⁺ complexes) were measured. These data allowed determining the stability constants by using the pH-potentiometric (by knowing the acid concentration of the samples and their pH at equilibrium) and ¹H relaxometric method (by using the pH and the relaxivities of the samples). These two sets of data were fitted simultaneously.

The equilibrium constants have been calculated from the ti-

tration data with the program PSEQUAD.⁶¹

Kinetic studies. The rates of the dissociation of the Ln³⁺ complexes in metal exchange reactions taking place with the Cu²⁺ ion were studied by spectrophotometry at 300 nm in thermostated cells (1.0 cm, 25 °C) with Cary 100 Bio (1,4-DO2APA³⁻) and Jasco V770 (1,7-DO2APA³⁻) spectrophotometers in the presence of 0.15 M NaCl. The concentration of the [Ln(1,4-DO2APA)] and [Ln(1,7-DO2APA)] complexes was 0.3 mM and 0.2 mM, respectively. The Cu²⁺ ion was used in 10- to 40-fold excess for the [Ln(1,4-DO2APA)] complexes, while a 10-fold excess was applied in the case of [Ln(1,7-DO2APA)] complexes. A non-coordinating buffer at 0.05 M concentration was used to keep the pH constant (DMP: 1,4-dimethylpiperazine, pK_a=4.17). The exchange reactions were studied in the pH range 3.4-5.0. For the calculation of the pseudo-first-order rate constants (*k_d*), the absorbance values measured at different *t* times were fitted to the following equation:

$$A_t = (A_0 - A_e)e^{-k_d t} + A_e \quad (7)$$

where A₀, A_t and A_e are the absorbance values measured at the start, at time *t* and at equilibrium, respectively.

Relaxometric measurements. Measurements of longitudinal relaxation times (*T*₁) were performed by using a Bruker Minispec MQ-20 NMR Analyzer operating at 0.49 T (corresponding to 20 MHz proton Larmor frequency) at 25.0±0.2 °C (set by a circulating water bath). The *T*₁ values were determined with the inversion recovery method (180° – τ – 90°) by averaging 4-6 data points obtained at 10 different τ delay values.

NMR studies. CEST spectra have been recorded at 25 and 37 °C, at diverse radiofrequency fields (B₁ = 2.5, 5, 10, 15, 20, 25 and 30 μT) on a Bruker Advance 300 MHz spectrometer. Z-spectra of 15 mM complex aqueous solutions were performed using a saturation time of 10 s and a 2 ppm frequency resolution. The pH of the solution was adjusted to pH 7.0 by adding 0.01 to 0.1 M NaOH or HCl solutions. The exact concentration of the solutions was determined by elemental analysis of the complexes.

DFT calculations. Full geometry optimizations of the [Ln(1,4-DO2APA)(H₂O)]·2H₂O, [Ln(1,7-DO2APA)(H₂O)]·2H₂O (Ln = La, Gd or Yb) and [Yb(1,4-DO2APA)] systems were performed employing DFT calculations at the TPSSh/LCRECP/6-31G(d,p)⁶² level with the Gaussian 09 package (Revision E.01).⁶³ LCRECP refers to the large-core quasi-relativistic pseudopotential of Dolg et al. and its associated (7s6p5d)/[5s4p3d]-GTO valence basis set.⁶⁴ Solvent effects were included by using the polarizable continuum model (PCM), in particular the integral equation formalism (IEFPCM) variant.⁶⁵ No symmetry constraints have been imposed during the optimizations. An ultrafine integration grid and the default SCF energy convergence threshold (10⁻⁸ a. u.) were used. The stationary points found on the potential energy surfaces as a result of geometry optimizations were characterized via frequency analysis. The electron density (ρ) and its Laplacian (∇²ρ) at the bond critical points (BCPs) were computed with the Multiwfn program (version 3.2).⁶⁶

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the [ACS Publications website](#) at DOI: XXXXX. Additional kinetic data, speciation diagrams, CEST spectra and DFT data (PDF).

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Notes

The authors declare no competing financial interest.

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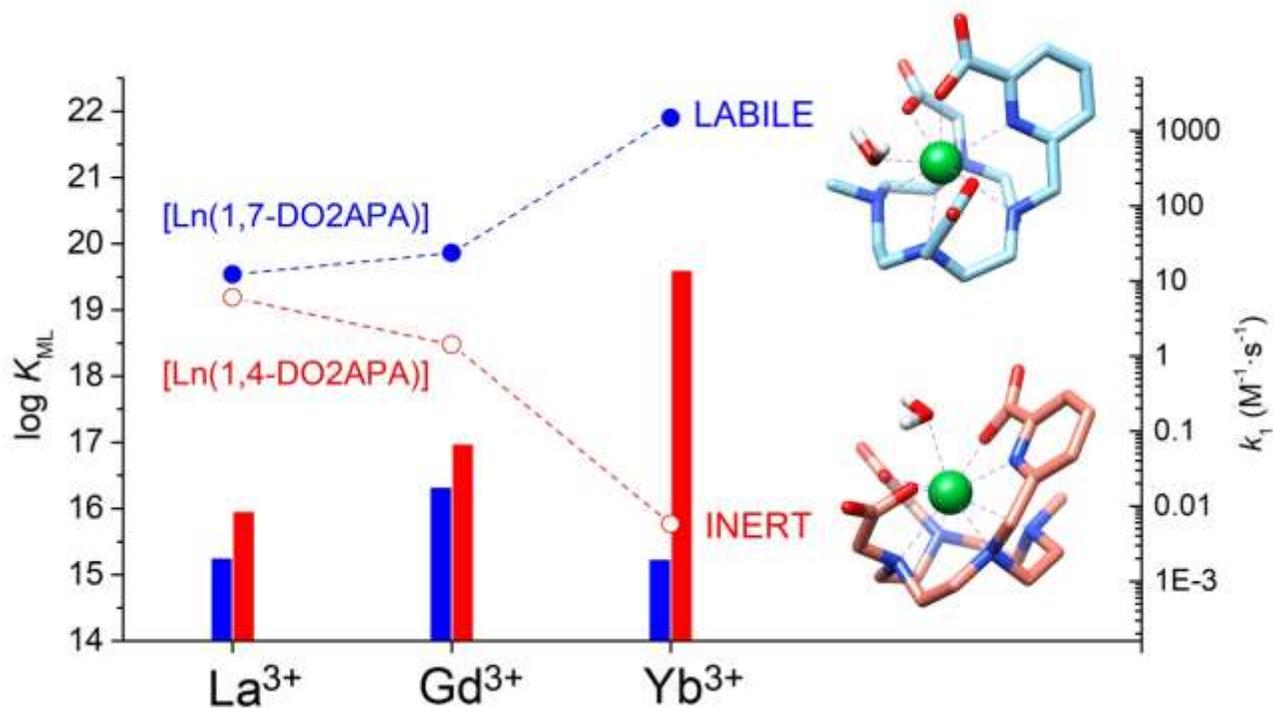
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TOC Graphic:



The arrangement of the ligand donor atoms around the lanthanide ion provokes dramatic differences in the thermodynamic stabilities and dissociation kinetics of lanthanide complexes, as demonstrated by investigating the complexes with two isomeric cyclen-based ligands, containing a picolinate arm and two acetate arms either in positions 1,7 or 1,4 of the macrocyclic structure.