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Amandine Conte-Daban, Maryline Beyler, Raphaël Tripier, Christelle Hureau. Kinetics are crucial when targeting copper ions to fight Alzheimer's disease: an illustration with azamacrocyclic ligands. *Chemistry - A European Journal*, 2018, 24 (33), pp.8447-8452. 10.1002/chem.201801520. hal-01764900

HAL Id: hal-01764900

<https://hal.univ-brest.fr/hal-01764900v1>

Submitted on 30 Apr 2018

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Kinetic are crucial when targeting copper ions to fight Alzheimer's disease: an illustration with azamacrocyclic ligands

Amandine Conte-Daban,^[a,b] Maryline Beyler,^[c] Raphaël Tripier^[c] and Christelle Hureau^{*[a,b]}

[a] Amandine Conte-Daban, Christelle Hureau
CNRS, LCC (Laboratoire de Chimie de Coordination)
205 route de Narbonne, BP 44099 31077 Toulouse Cedex 4, France
E-mail: christelle.hureau@lcc-toulouse.fr

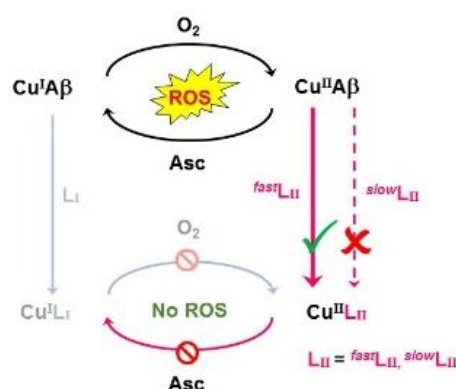
[b] Amandine Conte-Daban, Christelle Hureau
University of Toulouse
UPS, INPT, 31077 Toulouse Cedex 4, France

[c] Maryline Beyler, Raphaël Tripier
Université de Bretagne Occidentale
UMR-CNRS 6521 CEMCA, IBSAM
UFR des Sciences et Techniques, 6 avenue Victor le Gorgeu, C.S. 93837, 29238 BREST Cedex 3, France

Abstract: Targeting copper ions to either remove or redistribute them is currently viewed as a possible therapeutic strategy in the context of Alzheimer's disease (AD). Thermodynamic parameters, as for instance the copper(II) affinity of the drug candidate or the copper(II) over zinc(II) selectivity, are considered in the design of the drug candidate. In contrast, kinetic ones have been overlooked despite their probable high importance. In the present article, we use a series of azamacrocyclic ligands to demonstrate that kinetic issues must be taken into account when designing copper-targeting drug candidates in the context of AD.

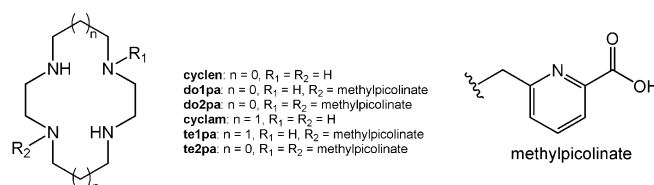
Introduction

AD is the most common neurodegenerative disorder characterized by the formation of extracellular senile plaques.^[1] They are found in AD brains and contain high levels of Cu^{2+} and Zn^{2+} ions embedded with aggregates of the amyloid- β ($\text{A}\beta$) peptide.^[2] Aggregates of $\text{A}\beta$ can be of various size and morphologies with distinct cell toxicity.^[3] While metal ions can effect aggregation pathways and final species, neither the mechanism nor the resulting type of aggregates is clear. Cu ions are essential and play key biological roles,^[4] but due to their ill-controlled redox ability they are harmful in AD.^[5] Actually, $\text{A}\beta$ -bound Cu ions in the presence of a reductant such as ascorbate (Asc) present at 200-400 μM in the synaptic cleft and of dioxygen, produce Reactive Oxygen Species (ROS) that participate in the oxidative stress observed in AD and damage neighboring biomolecules.^[5b] This is one of the main reason why molecules targeting copper ions, mainly Cu^{2+} , have been developed as drug candidates within the context of AD. Such ligands can remove Cu^{2+} from $\text{A}\beta$ (chelator) and optionally redistribute it intracellularly (metallophore) and many of them have been reported in the last years.^[6] They fulfill several requirements, including having a higher affinity for Cu^{2+} than the $\text{A}\beta$ peptide, the ability to redox silence the Cu^{2+} ion and a very high Cu^{2+} over Zn^{2+} selectivity.^[7] However, because the synaptic cleft where the Cu - $\text{A}\beta$ interaction occurs is proposed to be a quite reducing environment,^[8] Cu^{2+} might not be the most appropriate redox state of copper to target. Thus, a first possible strategy is to target Cu^+ ,^[9] or both redox states.^[10] A second one is to design Cu^{2+} drug candidates that are able to remove Cu^{2+} "fast enough" from the $\text{A}\beta$ peptide. Indeed during ROS production the copper ion switches between the +I and +II redox state; hence an efficient Cu^{2+} ligand should remove the transient Cu^{2+} from the $\text{A}\beta$ peptide and redox-silent it faster than Asc can reduce it (Scheme 1).



Scheme 1. Two approaches to stop Cu - $\text{A}\beta$ ROS production. On the one hand, Cu^{I} is targeted and the complex $\text{Cu}^{\text{I}}\text{L}_\text{I}$ is inert toward O_2 oxidation (shaded part of the scheme). On the other hand, Cu^{II} is targeted and the complex $\text{Cu}^{\text{II}}\text{L}_\text{II}$ is inert toward Asc reduction. Two possibilities occur: (i) if L_II is able to remove the Cu^{2+} from $\text{A}\beta$ faster (bold solid line, $\text{L}_\text{II}=\text{fastL}_\text{II}$) than Asc reduces it, then it is efficient (\checkmark); (ii) if not (dashed line, $\text{L}_\text{II}=\text{slowL}_\text{II}$) then it is inefficient (\times) although it could be able to remove Cu^{2+} from $\text{A}\beta$ on longer time scales in absence of Asc . \checkmark and \times refer to the notation used in Table 1.

Tetraazamacrocycles are ligands that fulfill the thermodynamic criteria stated above and are well known for the kinetic inertness and the thermodynamic stability of the complexes counterparts. In particular, 1,4,7,10-tetraazacyclododecane (cyclen) and 1,4,8,11-tetraazacyclotetradecane (cyclam) have been previously tested in the AD context with mitigating results.^[11] They are also the scaffolds of many chelators used for radiopharmaceutical purposes that bind $^{64/67}\text{Cu}^{2+}$, respectively for positron emission tomography and/or therapy and Gd^{3+} for magnetic resonance imaging, for instance.^[12] As a general trend, Cu^{2+} complexation with unsubstituted tetraazamacrocycles is quite slow and pending arms are added to improve the rate of Cu^{2+} complexation,^[13] as it was also observed with other cations.^[14] Inspired by these works, we report here the use of two families of tetraazamacrocycles based on the cyclen or cyclam scaffold as potential Cu^{2+} chelators against AD and illustrate the importance of kinetic issues in the removal of Cu^{2+} from the $\text{A}\beta$ peptide[†] and subsequent ROS arrest.



Scheme 2. Representation of the six different azaligands studied.

Results and Discussion

Scheme 2 shows the structure of the six ligands used in the present study. In addition to the non-functionalized cyclen and cyclam ligands, four macrocycles based on cyclen and cyclam ligands and bearing one (do1pa and te1pa, respectively) or two (do2pa and te2pa, respectively) methylpicolinate pendant arms were studied. Te2pa is new and as for its three analogues, has been obtained following the regiospecific N -functionalization of the macrocycle *via* the bisaminal chemistry (synthetic scheme S1). A view of the structure of the Cu^{2+} corresponding complex is shown in Figure 1, while bond distances and angles of the metal coordination

environment are given in the caption (see also Tables S1-S2). The central Cu^{2+} cation is six-coordinated in a highly elongated octahedral geometry, with the four nitrogen atoms of the macrocycle in the equatorial plane and the two nitrogen atoms of the picolinate arms at the apical positions. The Cu-NH (Cu-N₂) bond are shorter than the Cu-N₁ bond (1.9974(13) versus 2.0462(13) Å) whereas the apical Cu-N₃ bonds are much longer (2.7910(16) Å) due to the Jahn-Teller distortion. The carboxylate functions do not participate to the coordination of the metal center as observed for te1pa.^[15] This six-coordinated Cu^{2+} complex adopts the usually observed *trans*-III configuration (*R,R,S,S* conformation). The thermodynamic protonation constants of te2pa and its stability constants with Cu^{2+} were determined in 0.10 M KNO_3 at 25°C using potentiometric titrations.

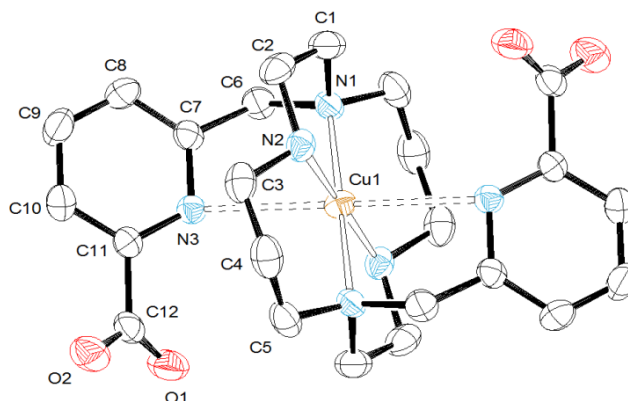


Figure 1. X-ray crystal structure of Cu^{2+} -te2pa with atom labelling; hydrogen atoms and water molecules are omitted for simplicity. The ORTEP plots is drawn at the 30% probability level. Cu-N(1)=2.0604(15), Cu-N(2)=1.9906(15), Cu-N(3)=2.7910(16), N(1)-Cu-N(2)=86.61(6), N(1)-Cu-N(2)=93.39(6), N(1)-Cu-N(3)=75.81(5), N(2)-Cu-N(3)=87.90(6), N(3)-Cu-N(1)=104.19. Symmetry transformations used to generate equivalent atoms: #1 $-x+1, -y+1, -z+1$.

The stepwise constants ($\log(K)$) and the overall constants ($\log(\beta)$) are collected in Table S3, while the speciation diagrams are showed in Figures S1-S2. The conditional affinity constant (corresponding to the absolute affinity constant at a given pH) has been calculated from the $\log(K)$ and $\log(\beta)$ values and equals $10^{14.9}$ at pH 7.1, a weaker value compared to affinity constants of the other five ligands under study (Table 1).^[15-16] XANES and EPR signatures of the six Cu^{2+} -L (L is any of the ligand studied) complexes in frozen solution are shown in Figures S3-S4 and corresponding EPR parameters are listed in Table 1. EPR signatures and parameters are consistent with an elongated octahedral geometry and a 4N binding set in the equatorial plane.^[17] XANES fingerprints of Cu^{2+} -te2pa is reminiscent of those of the parent Cu^{2+} -cyclam complex in line with the X-ray structure (Figure 1) that indicates an extremely long apical distance.

The cyclic voltammograms of the six complexes (Figure S5) indicate the possibility of reduction of the Cu^{2+} ion but at very low potentials (below -0.5 V versus SCE). As a general trend, addition of methylpicolinate arm(s) makes the reduction easier, i.e. the reduction occurs at higher potential (Table 1). Given the potential values of the cathodic process observed for all complexes, reduction by Asc is thermodynamically disfavored.^[10]

Table 1. Thermodynamic, spectroscopic and electrochemical characteristics of the six studied Cu^{2+} complexes.

	Cu^{2+} -cyclen	Cu^{2+} -do1pa	Cu^{2+} -do2pa	Cu^{2+} -cyclam	Cu^{2+} -te1pa	Cu^{2+} -te2pa
Coordination	4Neq 1Oap	5N 1O	N.D.[a]	4Neq	4 Neq 1Nap	4 Neq (2 Nap)
$\log K_a^{\text{Cu}}$ pH 7.1	17.4–18.8	18.5	17.4	19.9–20.8	18.0	14.9
refs.	[18]	[15]	[16]	[19]	[15]	This work
g_{\parallel}	2.20±0.01	2.23±0.01	2.23±0.01	2.18±0.01	2.19±0.01	2.18±0.01
$A_{\parallel}({}^{65}\text{Cu})$ (10^{-4} cm^{-1})	202±5	187±5	181±5	224±5	211±5	217±5
E_{pc} (V vs. SCE)	- 0.85	- 0.82	- 0.67	- 0.98	- 0.88	- 0.76
Reversibility ^[a]	Irrev.	Rev.	Rev.	Irrev.	Rev.	Irrev.
ROS	Incubation Cu or Cu-A β	✓ ^[c]	✓	✓	✓	✓
	No incubation Cu	~	✓	✓	×	✓
	Cu-A β	×	✓	~	×	✓

^[a] Not determined, ^[b] "Rev" indicates a electrochemically reversible or quasi-reversible process, "Irrev" indicates an electrochemically irreversible process, ^[c] ✓=total arrest, ×=no arrest, ~=incomplete arrest of ROS production.

In line with the electrochemical measurements, once the Cu^{2+} complexes are formed with any of the six ligands, they can't be reduced by Asc as probed by the absence of Asc consumption monitored by UV-Vis (Figure S6, Panels A and B). Similarly, Asc consumption was not observed when Cu^{2+} -A β 16 complex was pre-incubated with the ligand (Figure 2, Panels A and B). This is in line with the higher Cu^{2+} affinity values of the ligands compared to A β (Table 1) and the resistance to Asc reduction of the resulting Cu^{2+} -L complexes. Hence, the six macrocyclic ligands appear as good candidates for Cu^{2+} chelatotherapeutic purposes. In order to go further and challenge these ligands under more biologically relevant conditions, the Asc consumption assay was performed differently: Cu^{2+} (Figure S6, Panels C and D) or Cu^{2+} -A β (Figure 2, Panels C and D when A β = A β 16 and Figure S6, panels E and F when A β = A β 40) was first reacted with Asc and then the ligands were added. Such conditions better mimic the brain environment, which is at the same time rich in Asc and dioxygen.^[20]

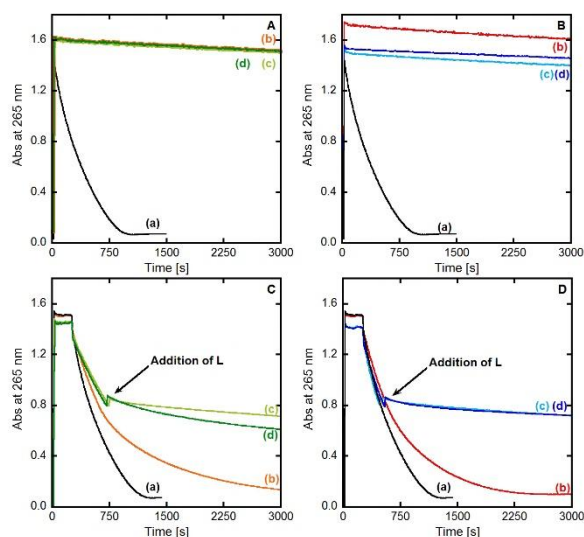


Figure 2. Kinetics of Asc consumption, followed by UV-visible spectroscopy at 265 nm with a background correction at 800 nm. *Panel A.* (a) A β 16+Cu $^{2+}$ +Asc, (b) A β 16+Cu $^{2+}$ +cyclen+Asc, (c) A β 16+Cu $^{2+}$ +do1pa+Asc, (d) A β 16+Cu $^{2+}$ +do2pa+Asc. *Panel B.* (a) A β 16+Cu $^{2+}$ +Asc, (b) A β 16+Cu $^{2+}$ +cyclam+Asc, (c) A β 16+Cu $^{2+}$ +te1pa+Asc, (d) A β 16+Cu $^{2+}$ +te2pa+Asc. *Panel C.* (a) Asc+A β 16+Cu $^{2+}$, (b) Asc+A β 16+Cu $^{2+}$ +cyclen, (c) Asc+A β 16+Cu $^{2+}$ +do1pa, (d) Asc+A β 16+Cu $^{2+}$ +do2pa. *Panel D.* (a) Asc+A β 16+Cu $^{2+}$, (b) Asc+A β 16+Cu $^{2+}$ +cyclam, (c) Asc+A β 16+Cu $^{2+}$ +te1pa, (d) Asc+A β 16+Cu $^{2+}$ +te2pa. [L]=[A β 16]=12 μ M, [Cu $^{2+}$]=10 μ M, [Asc]=10 μ M, [HEPES]=100mM, pH 7.1.

Under such conditions, the arrest of Asc consumption depends on the ligand: the arrest is total with do1pa, te1pa and te2pa, quasi-total with do2pa and absent for both cyclam and cyclen. It is worth noting the Asc consumption experiments mirror the formation of ROS, in particular H $_2$ O $_2$ and HO $^\bullet$,^[21] and is a robust way to document the ability of ligand in the removal and redox silencing of Cu-A β species.^[10, 22] Here, we also directly monitor the formation of HO $^\bullet$ using a fluorescent assay (Figure S7) and the results perfectly match those of the Asc consumption experiment. The different behavior of the six ligands with respect to the arrest of Asc consumption may be linked to the rate of Cu $^{2+}$ extraction from Cu $^{2+}$ -A β during the redox cycling. Indeed, in the course of the Asc consumption, there is a competition between Cu $^{2+}$ -A β reduction (and thus Asc consumption) and Cu $^{2+}$ extraction by the ligands. If the Cu $^{2+}$ extraction is not fast enough, the Cu $^{2+}$ -A β keeps on redox cycling and consuming Asc. Based on the properties of non-substituted and substituted ligands with respect to the rate of Cu $^{2+}$ complexation,^[13] it is anticipated that the addition of methylpicolinate arm(s) accelerates Cu $^{2+}$ removal from A β . Such assumption has been probed by the measurements of the rate of Cu $^{2+}$ binding to the six ligands without or with pre-incubation with the A β peptide (Figure 3, panels A- C and B-D, respectively). Several comments can be made: (i) for all the ligands, the kinetic of Cu $^{2+}$ removal from A β is slower than the Cu $^{2+}$ binding. This suggests that to remove Cu $^{2+}$ from A β , the ligand either acts on unbound Cu $^{2+}$ in pre-equilibrium with peptide-bound Cu $^{2+}$ or forms a ternary species with Cu $^{2+}$ -A β , the creation and/or destruction of which is the rate limiting step; (ii) in line with previous reports, cyclen binds Cu $^{2+}$ more rapidly than cyclam does,^[15] (iii) as anticipated, ligands with methylpicolinate arms bind Cu $^{2+}$ and remove it from A β faster than the unsubstituted parent ligands and (iv) within the series of methylpicolinate ligands, the trend is as follows: do2pa < te1pa \leq do1pa \leq te2pa (Figure S9 for the ordering of the three later ligands). This trend is in agreement with the ability of the ligand to stop Cu $^{2+}$ -A β catalyzed ROS formation (Table 1 and Figure 2).

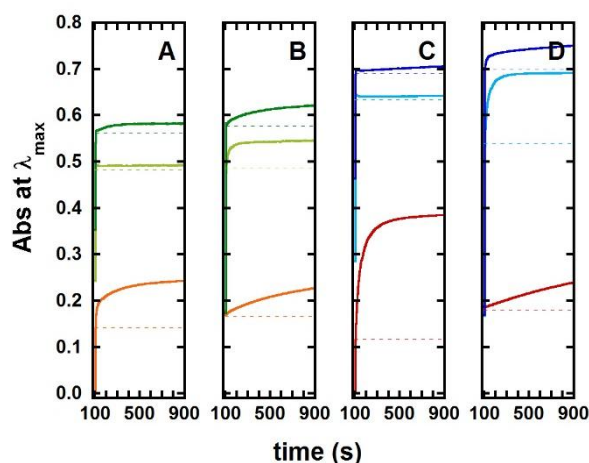


Figure 3. Kinetics of Cu^{2+} binding to the ligands (panels A and C) and after pre-incubation with $\text{A}\beta$ (panels B and D) at the wavelength of maximum intensity. Panel A: cyclen (orange), do1pa (light green) or do2pa (dark green) + Cu^{2+} ; Panel B: Cu^{2+} - $\text{A}\beta$ 16+cyclen (orange), do1pa (light green) or do2pa (dark green); Panel C: cyclam (red), te1pa (light blue) or te2pa (dark blue) + Cu^{2+} ; Panel D: Cu^{2+} - $\text{A}\beta$ 16+cyclam (red), te1pa (light blue) or te2pa (dark blue). $[\text{L}]=[\text{A}\beta$ 16] $=[\text{Cu}^{2+}]=50\mu\text{M}$, $[\text{HEPES}]=100\text{ mM}$, $\text{pH } 7.1$. Additions are performed at $t=100\text{s}$ and absorbance at the beginning of the kinetic is showed by a thin dotted horizontal line; see Figure S8 for the corresponding whole spectra plotted as a function of time.

As modulation of $\text{A}\beta$ aggregation by Cu^{2+} ions could potentiate the toxicity of the $\text{A}\beta$ aggregates *via* the formation of other morphological structures,^[22-23] we also test the six ligands for the ability to prevent formation of Cu^{2+} -type aggregates and to restore the formation of apo-like fibrils. All six ligands are able to do so, as evidenced by the monitoring of kinetic of fibrils formation using the classical ThT assay (Figure 4). More precisely, the formation of apo-type fibrils follows the sigmoid-type curve as expected for the nucleation – elongation – equilibrium three-step aggregation process. The $t_{1/2}$ value corresponding to the time at which the ThT fluorescence equals half of the maximum fluorescence, is about 20 hours. In contrast, in presence of Cu^{2+} ion, a weak increase of ThT fluorescence is observed in the first hours of the aggregation, followed by a stationary phase and then a sigmoid-like increase with a $t_{1/2}$ of about 60 hours. In presence of Cu^{2+} , the final ThT fluorescence intensity is about twice weaker than for the apo-peptide. Upon addition of any of the six ligands, the first weak increase of ThT fluorescence is lost and a pure sigmoid-like curve recovered with $t_{1/2}$ values spanning from 15 to 30 hours, a divergence from the $t_{1/2}$ value of the apo-peptide that is in the acceptable range of error bars of aggregation experiments. The kinetic data are confirmed by TEM (Figure 5). Apo-type fibrils consists in bold intertwined fibrils made of two to three protofilaments while the Cu^{2+} -type fibrils are longer and thinner mainly made of one unique protofilament. In presence of te2pa on the $\text{A}\beta$ 40 + Cu^{2+} sample, the morphology of apo-type fibrils is recovered. The identical activity of the six ligands observed here is linked to the time scale of the aggregation experiments (several hours compared to the seconds time scale of the ROS experiments), which allows Cu^{2+} to be removed from the $\text{A}\beta$ peptide by any of them.

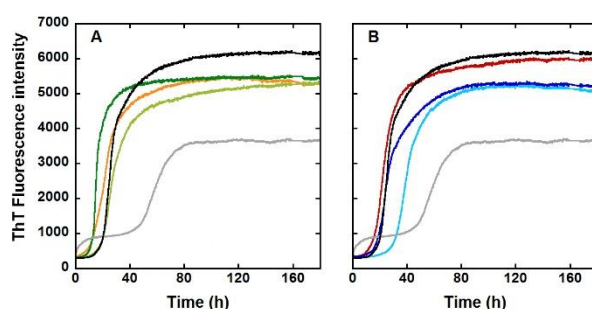


Figure 4. Aggregation curves of $\text{A}\beta$ peptide. *Panel A.* Kinetics of ThT fluorescence of $\text{A}\beta$ (black curve), Cu^{2+} - $\text{A}\beta$ (grey curve), Cu^{2+} - $\text{A}\beta$ + cyclen (orange curve), Cu^{2+} - $\text{A}\beta$ + do1pa (light green curve), Cu^{2+} - $\text{A}\beta$ + do2pa (dark green curve). *Panel B.* Kinetics of ThT fluorescence of $\text{A}\beta$ (black curve), Cu^{2+} - $\text{A}\beta$ (grey curve), Cu^{2+} - $\text{A}\beta$ + cyclam (red curve), Cu^{2+} - $\text{A}\beta$ + te1pa (light blue curve), Cu^{2+} - $\text{A}\beta$ + te2pa (dark blue curve). $[\text{L}]=[\text{A}\beta$ 40] $=20\mu\text{M}$, $[\text{Cu}^{2+}]=18\mu\text{M}$, $[\text{ThT}]=10\mu\text{M}$, $[\text{HEPES}]=50\text{ mM}$, $\text{pH } 7.1$, $T=37^\circ\text{C}$.

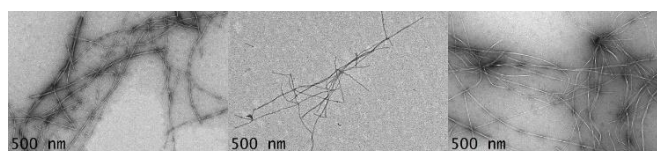


Figure 5. TEM images of $\text{A}\beta$ 40 (left image), Cu^{2+} - $\text{A}\beta$ (central image) and $\text{A}\beta$ 40 + Cu^{2+} + te2pa (right image).

Conclusions

In the present article, we report the study of six macrocyclic ligands to demonstrate that kinetic issues are important in the design of new chelator/metallophore in the context of AD. The study was performed with the A β peptide starting at position 1, considered as the most deleterious with respect to ROS production. More recently, indeed, N-terminally truncated peptides starting at position 4 were reported to be present in significant amount in the synaptic cleft and amyloid plaques,^[24] but their Cu²⁺ complexes showed much more less propensity to produce ROS.^[25]

Here the objective is to influence the future design of drug candidates against AD and other diseases involving Cu-induced oxidative stress, such as Parkinson disease.^[26] We show that kinetic parameters can be as important as thermodynamic or structural ones. This is particularly obvious when comparing the te2pa and the parent cyclam ligand: while the Cu²⁺ environments are very close in both Cu²⁺ complexes (the two picolinate arms are indeed very far from the Cu²⁺ center in the Cu²⁺-te2pa complex) and the affinity of the latter is higher by about 5 orders of magnitude, the te2pa can remove Cu²⁺ faster from A β and is active against ROS production in contrast to the parent cyclam. No general trend can be drawn within the family of azamacrocyclic ligands, since for instance in the cyclen series, the addition of a sole picolinate arm leads to a faster Cu(II) removal from A β than with the addition of two picolinate arms while in the cyclam series, this is the opposite. To draw a structure – activity relationship the series of ligands studied has to be enlarged.

More generally, having ligands able to remove copper from A β fast enough may be one overlooked requirement to take into account in the design of drug-candidates targeting Cu ions to fight AD. It adds to other better-known requirements such as being able to cross the blood brain barrier, having Cu²⁺ affinity higher than A β , not removing Cu²⁺ ions from essential enzymes and being selective for Cu²⁺ against Zn²⁺.^[27] The kinetic counterpart of the latter point, that is to say, the effect of Zn²⁺ ions on kinetic processes related to AD is currently under study and will be highlighted in further forthcoming works.

In a period where the deleterious intervention of metal ions and A β peptides in AD are disputed,^[28] due to the failure of the first generation of metallophores, one might wonder whether targeting copper ions is still a viable approach. With the present study, we show that an alternative way of questioning the metal hypothesis is to ask ourselves whether all requirements have been taken into account in the design of drug candidates targeting metal ions.

Experimental Section

Synthesis. Reagents were purchased from ACROS Organics and from ALDRICH Chemical Co. Cyclam was purchased from Chematech (Dijon, France). Bisformyl-Cyclam (**1**)^[29] and methyl 6-(chloromethyl)picolinate (**2**)^[30] were synthesized as previously described (See Scheme S1). The solvents were freshly distilled prior to use and according to the standard methods. NMR spectra (¹H and ¹³C) were recorded at the core facilities of the University of Brest, with Bruker Avance 500 (500 MHz) or Bruker AMX-3 300 (300 MHz) spectrometers. The HR-MS analyses were performed at the Institute of Analytic and Organic Chemistry, ICOA in Orléans. Reagents, except the different ligands, were commercially available and were used as received. All the solutions were prepared in milliQ water (resistance: 18.2 M Ω .cm).

The Cu²⁺ ion source was CuSO₄.5H₂O, bought from Sigma-Aldrich.

HEPES buffer (sodium salt of 2-[4-(2-hydroxyethyl)piperazin-1-yl]ethanesulfonic acid) was bought from Sigma-Aldrich. A stock solution was prepared at 500 mM, pH = 7.1.

Phosphate buffer was bought from Sigma-Aldrich. Two stock solutions, K₂HPO₄ and KH₂PO₄, were prepared at 500 mM, and they were mixed until to reach a stock solution at pH = 7.1.

Sodium ascorbate was bought from Sigma-Aldrich. A stock solution was prepared at 5 mM each day because of the quick degradation of the ascorbate.

Thioflavin T (ThT) was bought from Acros Organics. A stock solution of ThT at 250 μ M was prepared in water without any further purification.

Peptides. A β 16 (DAEFRHDSGYEVHHQK) was bought from Genecust. A stock solution of about 10 mM was prepared and titrated using the Tyr chromophore, with $\epsilon = 1410 \text{ cm}^{-1} \text{ M}^{-1}$ at acidic pH. The stock solution was stored at 4°C. A β 16 is considered as a good model of the A β 40 with respect to metal ions binding.^[5a] A β 40 (DAEFRHDSGYEVHHQKLVFFAEDVGSNKGAIIGLMVGGVV) was bought from Genecust. Around 6 mg were dissolved in approx. 400 μ L of NaOH 50 mM. This solution was purified by FPLC, with a Superdex 75 column and NaOH 15 mM as eluent, with a flow rate at 0.9 mL min⁻¹. The collected fractions were titrated using the Tyr chromophore, with $\epsilon = 2400 \text{ cm}^{-1} \text{ M}^{-1}$ at basic pH. The stock solution was directly used for the ThT experiments.

Single crystal X-ray diffraction measurements. Single-crystal X-ray diffraction data were collected at 170 K on an X-CALIBUR-2 CCD 4-circle diffractometer (Oxford Diffraction) with graphite-monochromatized MoK α radiation ($\lambda = 0.71073$). Crystal data and structure refinement details are given in Table 5. Unit-cell determination and data reduction, including interframe scaling, Lorentz, polarization, empirical absorption and detector sensitivity corrections, were carried out using attached programs of CrysAlis software (Oxford Diffraction).^[31] Structures were solved by direct methods and refined by full matrix least squares method on F2 with the SHELXL^[32] suites of programs. The hydrogen atoms were identified at the last step and refined under geometrical restraints and isotropic U-constraints.^[33] CCDC number 1540075 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

UV-Visible spectrophotometry. UV-Vis ROS kinetic experiments were recorded on a spectrophotometer Agilent 8453 at 25°C in 1 cm path length quartz cuvette, with an 800 rpm stirring. The samples were prepared from stock solutions of ligand, peptide and Cu²⁺ diluted down to 12, 12 and 10 μ M respectively in 100mM HEPES solution, pH = 7.1. Ascorbate is diluted down to 100 μ M. UV-Vis kinetic monitoring of the Cu²⁺ binding or Cu²⁺ removal from A β by the six ligands were recorded on a spectrophotometer Agilent 8453 at 25°C in 1 cm path length quartz cuvette, with an 800 rpm stirring. The samples were prepared from stock solutions of Cu²⁺, peptide and ligand diluted down to 50, 50 and 50 μ M respectively in 100mM HEPES solution, pH = 7.1.

Fluorescence experiments. ThT experiments were recorded on a FLUOstar OPTIMA BMG LABTECH at 37°C in a 384-well plate bought from Dutscher SAS. ThT was excited at 440 nm and the fluorescence was recorded at 490 nm. The gain was 1200. The samples were prepared from stock

solutions of ligand, peptide and Cu²⁺ diluted down to 20, 20 and 18 μM respectively in 100 mM HEPES buffer, pH = 7.1. ThT was added at a resulting concentration of 10 μM .

Transmission electron microscopy. Specimens were prepared for electron microscopy using the conventional negative staining procedure. 20 μL of solution was absorbed on Formvar-carbon-coated grids for 2 min, blotted, and negatively stained with uranyl acetate (1%) for 1 min. Grids were examined with a TEM (Jeol JEM-1400, JEOL Inc, Peabody, MA, USA) at 80 kV. Images were acquired using a digital camera (Gatan Orius, Gatan Inc, Pleasanton, CA, USA) at a x 25 000 magnification.

Supporting information. Details of potentiometry measurements and analysis, of EPR, UV-Vis, XANES and Fluorescence spectroscopy and TEM experiments are given in the Supporting Information.

Acknowledgements

The authors acknowledge the ESRF Synchrotron, proposal CRG-FAME-BM30B 30-02-1100 for providing beam time and for their help in recording the XANES spectra.

Keywords: copper • peptide • bioinorganic • chelator • azamacrocycles

Abbreviations : A β , Amyloid- β ; Abs, absorbance; AD, Alzheimer's disease; Asc, ascorbate; EPR: Electron Paramagnetic resonance; XANES (X-ray Absorption Near-Edge Structure) ; HEPES, 4-(2-hydroxyethyl)piperazine-1-ethanesulfonic acid; ROS, Reactive Oxygen Species ; ThT : Thioflavine T ; TEM: Transmission Electronic Microscopy.

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[†] Note that in the text A β refers to the C-terminally truncated A β 16 peptide regarded as an appropriate model of the full-length A β peptide with respect to metal ions binding and/or to the full-length A β 40 peptide, unless when specific label is required for clarity purposes.

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